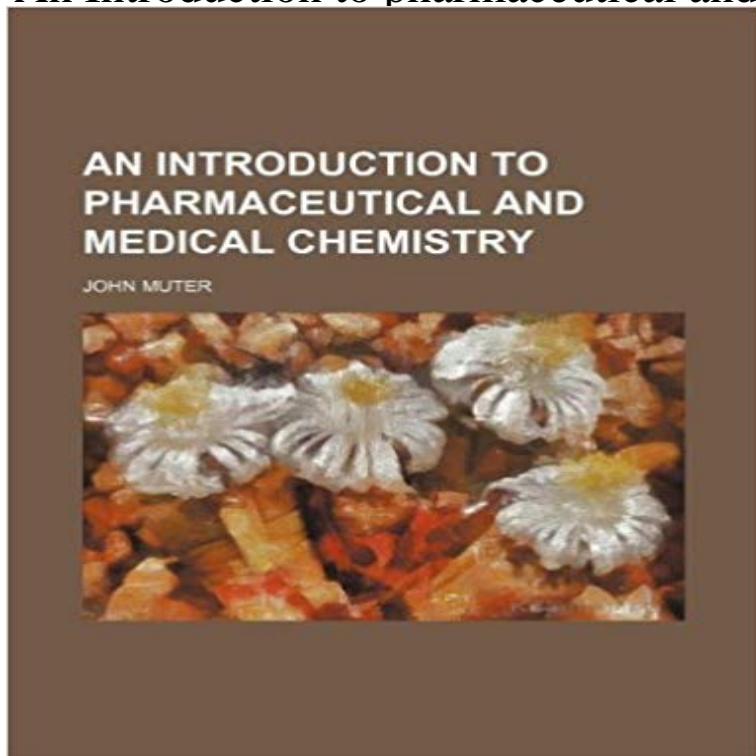


An Introduction to pharmaceutical and medical chemistry



This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1880 Excerpt: ...(Propene chloride.) METHENE CHLORIDE (syns. Methylene Chloride. Didilorometlianc). Formula, CH₂Cl. According to Perkin, this body is best obtained by agitating an alcoholic solution of chloroform with powdered metallic zinc and a small quantity of ammonium hydrate. The reaction is doubtless brought about by nascent hydrogen, and is accompanied by the formation of marsh-gas:--2CHCl₃ + 4H₂ = CH₂ + CH₄ + 4HCl. It is a colourless liquid, boiling at 106 F., and not solidifying at 0 F. It has been used as an anaesthetic under the name of bichloride of methylene. ETHENE CHLORIDE (syns. Ethylene Chloride. Dutch Liquid). Formula, C₂H₄Cl₂. Is produced by causing defiant gas and chlorine to mix together over water, when a yellowish oily liquid forms and sinks to the bottom in drops. It possesses a peculiarly fragrant odour, and burns with a greenish flame. It is interesting as being one of the early discoveries in this branch of chemistry, having been first produced by four Dutch chemists in 1795. (c) Triatomic Haloid Ethers, Series CaH_jn.jX_j. METHENYL CHLORIDE (syns. Chloroform. Formyl Chloride). Formula, CHCl₃. Is prepared by distilling spirit of wine with chlorinated lime and slaked lime in the presence of a considerable quantity of water, when chloroform and calcium formate are produced. This may be represented in one equation as follows:--4C₂H₅HO + 8Ca₂CO₃ = 2CHCl₃ + 3Ca₂CO₃ + 5CaCl₂ + 8H₂O. The real reactions which occur are more probably as follows. Keeping in mind that chlorinated lime can exercise both an oxidizing and a chloridizing influence, we may consider

that by the former the alcohol is reduced to aldehyd, and that by the latter hydrogen is displaced, forming trichloraldehyd, or chloral:-- $2\text{C}_2\text{H}_5\text{HO} + 6\text{Cl}_2 + \text{O}_3 = 2\text{aHC13O} + 6\text{HC1} + 2\text{O}_2$

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